



HELLENIC REPUBLIC

UNIVERSITY OF CRETE

Academic English

Section: Intermolecular forces Worksheet

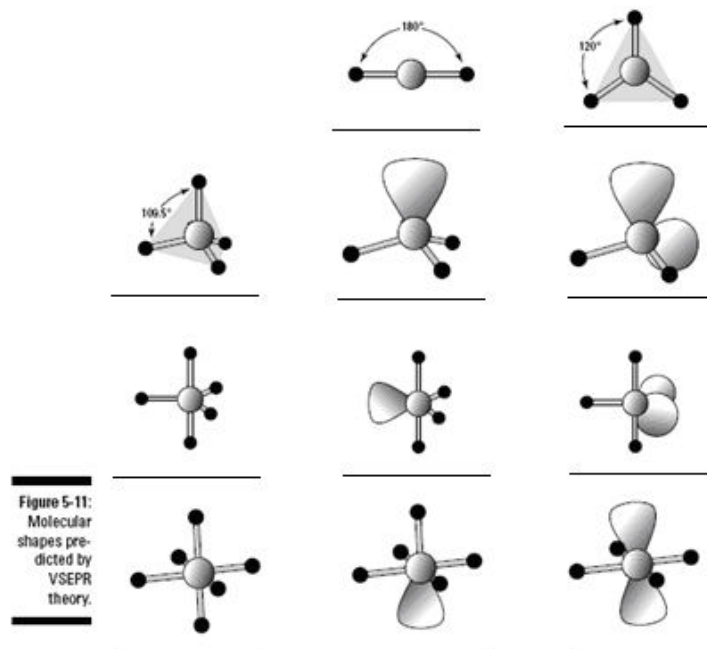
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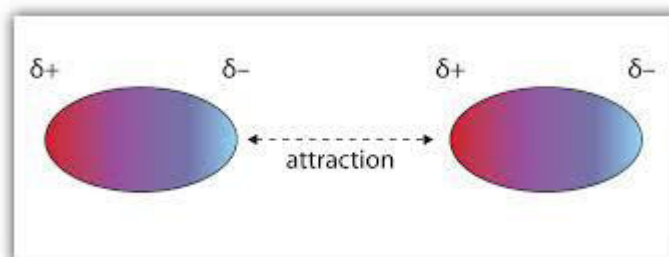
INTERMOLECULAR FORCES WORKSHEET

TASK: Describe the shapes of molecules using the terms given:

Tetrahedral, octahedral, linear, bent (V shaped), trigonal pyramidal, trigonal planar, seesaw, square pyramidal, T shaped, square planar.



Types of Intermolecular forces



Task 2: *Watch the YouTube video on Intermolecular Forces (IMFs) and complete the missing information: <https://www.youtube.com/watch?v=90q7xl3ndJ8>*

Hydrogen bonds are not really bonds. They are called bonds **because**.....

Hydrogen Bond Definition: When a hydrogen molecule is attracted to a

Dipole-Dipole occurs when we have a polar bond. A polar bond occurs when

London Dispersion Forces involve

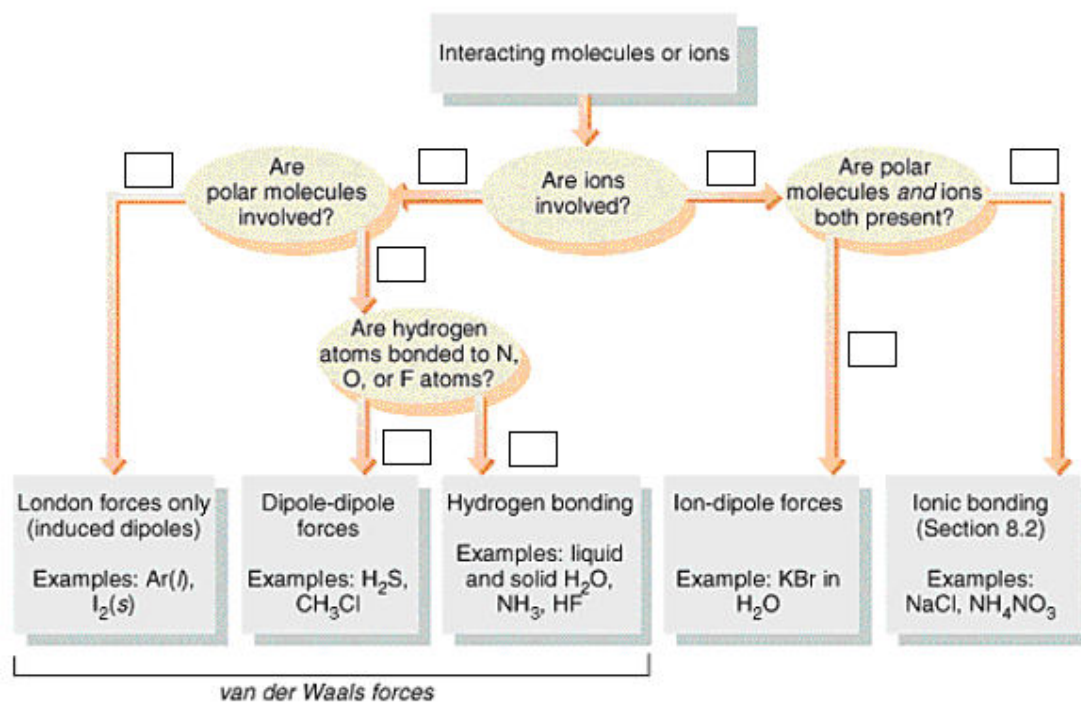
Put the above IMFs in the right order in terms of strength, starting with the strongest one.

- 1.
- 2.
- 3.

What does δ^+ and δ^- denote?

What's the difference between intramolecular and intermolecular forces?

TASK 3: Complete the boxes with Yes or No.



Notes

Reference Note

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